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**Medical**

# DHORT

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**Question:** 97

What property of a metal refers to its ability to be hammered into sheets?

- A. Ductility
- B. Thermal conductivity
- C. Electrical conductivity
- D. Malleability
- E. Density

**Answer:** D

Malleability refers to a metal's ability to be hammered into sheets.

**Question:** 98

Which of the following elements is most easily oxidized?

- A. Nitrogen
- B. Fluorine
- C. Lithium
- D. Neon
- E. Sulfur

**Answer:** C

Easily oxidized elements readily release electrons. The loss of an electron allows these elements to form a stable valence electron configuration. In the context of this question, lithium is the most willing to release an electron, so it is the most easily oxidized.

**Question:** 99

The hamstrings are responsible for flexing which joint?

- A. Knee
- B. Hip
- C. Shoulder

D. Elbow

**Answer:** A

The hamstrings are responsible for flexing the knee joint. The hamstrings consist of biceps femoris, semimembranosus and semitendinosus. They also assist in knee rotation.

**Question:** 100

What type of bonding does a molecule of NaCl display?

- A. Ionic
- B. Polar covalent
- C. Nonpolar covalent
- D. Metallic
- E. Covalent metallic

**Answer:** A

NaCl is composed of two ions,  $\text{Na}^+$  and  $\text{Cl}^-$ . Therefore, the molecule displays ionic bonding.

**Question:** 101

Which of the following compounds is classified as a metallic oxide?

- A.  $\text{H}_2\text{O}$
- B.  $\text{Na}_2\text{O}$
- C.  $\text{CO}_2$
- D.  $\text{NO}_2$
- E.  $\text{SO}_2$

**Answer:** B

Metallic oxides consist of an oxygen atom bound to a metal. The only metal present in this question is sodium, so  $\text{Na}_2\text{O}$  must be a metallic oxide.

**Question:** 102

A saturated solution of NaCl is heated until more solute can be dissolved. How is this solution best described?

- A. Dilute
- B. Immiscible
- C. Supersaturated
- D. Unsaturated
- E. Hydrophobic

**Answer:** C

When a saturated solution is heated so more solute can be dissolved, the solution is described as supersaturated. Supersaturated solutions are typically unstable, and the solute can crash out of solution if a seed crystal is provided.

**Question:** 103

A 40.0 gram sample of a radioactive element decays to 5.0 grams in 15 hours. What is the half-life of this element?

- A. 3 hours
- B. 2 hours
- C. 6 hours
- D. 7.5 hours
- E. 5 hours

**Answer:** E

In 15 hours, this substance has decayed to  $1/8$  of its original mass. In other words, the substance has progressed through three half-lives ( $1/2 \times 1/2 \times 1/2 = 1/8$ ). Thus, a single-half life for this substance is 5 hours.

**Question:** 104

Which of the following compounds contains a double bond?

- A.  $C_2H_6$
- B.  $C_2H_4$
- C.  $C_3H_8$
- D.  $CH_4$
- E.  $C_4H_{10}$

**Answer:** B

Hydrocarbons with the formula  $C_nH_{2n}$  contain double bonds. The only compound with this formula,  $C_2H_4$ , must contain a double bond.

**Question:** 105

What is the duodenum responsible for?

- A. Breaking down food in the small intestines
- B. Cleans the blood
- C. Creates bile
- D. Absorbs oxygen

**Answer:** A

The duodenum is responsible for breaking down food in the small intestines.

**Question:** 106

80.0 grams of NaOH is dissolved in 3.0 moles of  $H_2O$ . What is the mole fraction of NaOH in this solution?

- A. .20
- B. .37
- C. .40
- D. .64
- E. .79

**Answer:** C

The mole fraction for a compound indicates (moles of given compound) / (total moles of system). In this question, there are two moles of NaOH and five total moles in the system. Thus, the mole fraction is expressed as  $2/5$ , or .40.



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